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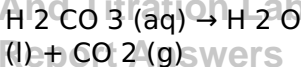
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## **Antacid Analysis And Titration Lab**

The calculated value of the base can then be used in the “back-titration” of another reagent as is done in part B where an antacid, composed mainly of  $\text{CaCO}_3$  (s), is used to neutralize the acidic  $\text{HCl}$  (aq) solution. The reaction is as follows:  $\text{CaCO}_3$  (s) +  $2\text{H}^+$  (aq)  $\rightarrow$   $\text{Ca}^{2+}$  (aq) +  $2\text{H}_2\text{O}$  (l) +

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**Acid-Base Titrations:  
Standardization of  
NaOH and Antacid**

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## **Antacid Analysis and Titration - Hands-On Labs**

1 Experiment 7:  
Titration of an Antacid  
Objective: In this experiment, you will standardize a solution of base using the analytical technique known as titration. Using this standardized solution,

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## Antacid Analysis

### And Titration Lab

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you will determine the acid neutralizing power of a commercially available antacid tablet.

### **Experiment 7: Titration of an Antacid**

titration with NaOH to figure out the amount of excess acid. Then, from this, we can calculate how much acid reacted with the antacid. This method of analysis is called a



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back-titration. The reactions above are reversible, which means that  $\text{CO}_2$  dissolved in water will produce some carbonic acid.

## **Titration of a Commercial Antacid**

Unformatted text  
preview: Antacid  
Analysis and Titration  
— Lab Report Assistant  
Exercise 1: Back  
Titration of Antacid  
Neutralization Data

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## Antacid Analysis And Titration Lab Report Answers

Table 1. Antacid  
Neutralization Data  
Mass of Crushed  
Antacid (g)  
Concentration of HCl  
(M) 1M — . a. Data  
Table 2.

### **Antacid Analysis and Titration - Antacid Analysis and ...**

Antacid Analysis and  
Titration-Lab Report  
Assistant Exercise 1:  
Back Titration of  
Antacid Neutralization  
Data Table 1. Antacid

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Neutralization Data  
Mass of Crushed  
Antacid (g)  
Concentration of HCl  
(M) Volume HCl (mL)  
Concentration of NaOH  
(M) Initial NaOH  
Volume (mL) Final  
NaOH Volume (mL)  
Total Volume of NaOH  
Used (mL) 1M 5 mL 1M  
Data Table 2.

**Solved: Antacid  
Analysis And  
Titration-Lab Report  
Assistan ...**

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## Antacid Analysis And Titration Lab Report Answers

This experiment was performed to learn the technique of acid-base titration and to compare the efficiency of commercially available antacids by looking at their weight of HCl and weight of antacid values. The analysis of antacid tablets was highlighted in this experiment. The efficiency of antacid tablets was determined and compared when the ..

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## **Acid-Base Titrations: Analysis of Antacid Tablets | Essay ...**

Lab 4 - Determination of the Amount of Acid Neutralized by an Antacid Tablet Using Back Titration Goal and Overview Antacids are bases that react stoichiometrically with acid. The number of moles of acid that can be neutralized by a single tablet of a commercial antacid will

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be determined by back titration. To do the experiment, an antacid tablet will be dissolved in a known excess amount of acid.

## **Lab 4 - Determination of the Amount of Acid Neutralized by ...**

This is from the Antacid Analysis and Titration.

Thank you!!! A.)

Calculate the number of moles of NaOH used to neutralize the 10mL

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of 1.0M HCl. B.)

Calculate the amount  
of HCl (grams)  
neutralized by the 1.0M  
NaOH in the control  
experiment. C.)

Calculate the  
difference in NaOH  
volume between the  
back titration average  
and the control  
experiment ...

## **Solved: This Is From The Antacid Analysis And Titration. T ...**

1. Weigh each tablet

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on the electronic scale.

2. Crush one of each tablet into powder using mortar and pestle. 3. Measure about 20 mL of HCl in a graduated cylinder. Move it to an erlenmeyer flask. 4. Measure 0.3 grams of a given antacid on a tared piece of paper on an electronic

**Analysis of Antacid  
Tablets by rebekeh  
ferrini on Prezi Next**



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In this experiment, a Back-Titration technique is used to determine the amount of acid neutralized by two different brands of antacid tablets. The back-titration is performed by adding a measured excess of standardized acid to a weighed sample of an antacid tablet. Some of the acid is neutralized by the tablet and some will remain.

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**Antacid Analysis: A  
Back-Titration**  
Report Answers

Abstract: Antacid analysis and titration may be used in order to determine the effectiveness of an antacid tablet, TUMS, in neutralizing acid. Testing the effectiveness of the tablet will be done with a technique, back titration, which is similar to standard titration. By adding acid to the tablet early

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on in the procedure the  
addition of NaOH, the  
titrant, can still be  
added in order to ...

**Lab 6 - Experiment 6**  
**Antacid Analysis**  
**Titration Name ...**

Then using the formula  
moles of HCl- moles of  
NaOH= moles of base  
in the antacid, there  
will be a lower number  
moles of the base in  
the antacid. Part B.1.  
An air bubble was  
initially trapped in the

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## Antacid Analysis And Titration Lab

buret but was dispensed during the back titration of the unreacted HCl (Part B.2).

### **Experiment 17 Post Lab: Antacid Analysis Flashcards | Quizlet**

Tared scale with 100-mL glass beaker.  
[www.HOLscience.com](http://www.HOLscience.com)  
13 ©Hands-On Labs, Inc. fExperiment Antacid Analysis and Titration 17. Use the

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## Antacid Analysis

### And Titration Lab

clean spoon to carefully weigh 0.5 g of the crushed antacid tablet into the glass beaker on the tared scale. Record this mass of the crushed antacid in Data Table 1 in your Lab Report Assistant.

## **42-0139-00-02-EXP, Antacid Analysis and Titration (1 ...**

TITRATION:

STANDARDIZATION OF  
A BASE AND ANALYSIS  
OF STOMACH ANTACID

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TABLETS Teacher

Notes This experiment is designed for students working singly or in groups of two. The overall purpose of this experiment is to determine the effectiveness of two different brands of stomach antacid tablets.

### **Analysis of stomach antacids teacher notes**

The titration equation

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can also be expressed as  $\text{HCl(aq)} + \text{NaCl} + \text{H}_2\text{O}$ . In order to find out which antacid worked the best, we started out with measuring 2g of our antacid and mixing it with 25mL of our HCl. These solutions were heated gently to a boil with a Bunsen burner for about 2min.

**Experiment 17 Lab  
report chem 112 -  
StuDocu**

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The object of this laboratory activity is to become familiar with making solutions and to titrate an acid with a base. One solution will be prepared from a solid and one solution will be prepared by dilution of a concentrated solution.

## **Analysis of stomach antacids - chymist.com**

This video describes how to calculate the



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cost effectiveness of an antacid tablet by using data gathered from a titration experiment. Detailed discussion on experimental set up is also provided.

### **Antacid Lab**

Antacid IV - 2 the first mechanism. The amount of acid neutralized will be measured through a process known as back titration. This is done by adding a known

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volume and concentration of HCl to the antacid, allowing it to react, and then using a known concentration of NaOH to bring the solution back to a neutral solution.

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